

Balancing Chemical Equations (Pg. 131-132)

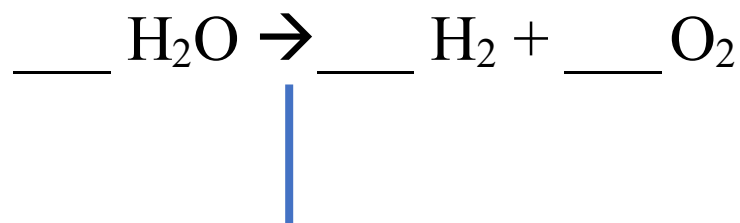
What is the law of conservation of mass and how is it related to balancing chemical equations?

Label this equation with “product side” and “reactant side”:



Before balancing the reaction, how many oxygen atoms are there on the reactant side and how many are on the product side?

Balance this reaction by drawing diagrams. What do the big numbers mean?



Synthesis Reactions (Pg. 156)

What is a synthesis reaction?

Draw a visual representation of a synthesis reaction:

Summarize 1 of the synthesis reactions given in the textbook

Burning Magnesium Ribbon

Materials: magnesium ribbon
tubing
bunsen burner
sparker
crucible tongs

Procedure:

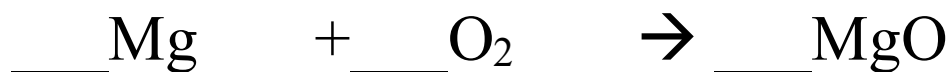
Note: DO NOT LOOK DIRECTLY AT THE BURNING RIBBON.

1. Connect your hose and Bunsen burner to the gas supply
2. Turn the gas supply on.
3. Spark the gas above the Bunsen burner to create a controlled burn
4. Hold the magnesium into the flame.
5. Record your observations.

Observations? What happened?

What kind of energy is released during the formation of magnesium oxide?

Balance this Chemical Equation:



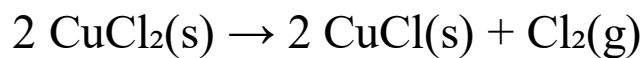
Diagrams

Diagrams

Sprinkle copper (II) chloride into a flame.

Observations? What happened?

Diagrams



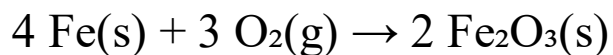
Diagrams



Sprinkle iron filings into a flame

Observations? What happened?

Diagrams



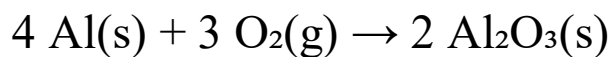
Diagrams



Sprinkle powdered aluminum into a flame

Observations? What happened?

Diagrams



Diagrams



Balance these Formation/Decomposition Reactions.

