

Breaking Bonds & Rearranging Atoms (Pg. 123-124)

What happens to chemical bonds during a chemical reaction?

How is energy involved in the breaking and formation of chemical bonds?

What determines the amount of energy needed to break a chemical bond?

How does the formation of chemical bonds contribute to the stability of atoms?

What happens to the energy when chemical bonds are formed?

How do chemical reactions affect the valence shells of atoms?

IONIC COMPOUNDS WITH POLYATOMIC IONS

Name: magnesium sulfate

Solid Description: _____

Aqueous Description: _____

Name: potassium permanganate

Solid Description: _____

Aqueous Description: _____

Name: potassium chromate

Solid Description: _____

Aqueous Description: _____

Formula: CaCO₃

Solid Description: _____

Aqueous Description: _____

Formula: CuSO₄

Solid Description: _____

Aqueous Description: _____

Formula: K₂Cr₂O₇

Solid Description: _____

Aqueous Description: _____

Diatomic Elements and Polyatomic Ions (Pg. 130)

Which elements are mentioned as diatomic in the text?

How are subscripts in chemical formulas different from coefficients in chemical equations?

What is the exception among the diatomic elements concerning their position on the periodic table?

Where are coefficients typically written in chemical equations?

How is the number of atoms in a single molecule or formula unit indicated in a chemical formula?

IONIC COMPOUNDS WITH POLYATOMIC IONS

Formula: BaCO_3

Name:

Bohr Diagram:

Lewis Dot Diagram

Formula: NH_4OH

Name:

Bohr Diagram:

Lewis Dot Diagram

Formula: NaNO_3

Name:

Bohr Diagram:

Lewis Dot Diagram

Name: ammonium hypochlorite

Name:

Bohr Diagram:

Lewis Dot Diagram

Name: sodium sulphate

Name:

Bohr Diagram:

Lewis Dot Diagram

Name: iron (III) hydroxide

Name:

Bohr Diagram:

Lewis Dot Diagram